



برعاية معالي وزير التربية والتعليم

السيد الأستاذ / محمد عبد اللطيف

وتوجيهات مساعد الوزير لشئون تطوير المناهج التعليمية
والمشرف على الإدارة المركزية لتطوير المناهج
د/ أكرم حسن

WEEK

(5)



الصف الثالث الإعدادي

Home work

الفصل
الدراسي الثاني

لجنة الأعداد والمراجعة
خبراء مكتب تنمية مادة العلوم

إشراف علمي
مستشار العلوم
د / عزيزة رجب خليفة

Lesson Two:Speed of Chemical ReactionsPart 2Question One:Choose the correct answer

1- The number of collisions between molecules increases and so the rate of the chemical reaction increases

- A- Increase in temperature.
- B- Increase in concentration of products.
- C- Increase in concentration of reactants.
- D- A, C together.

2- All of the following are from the characteristics of the catalyst except

- A- Its mass decreases
- B- It binds to the reactants during the reaction and then separates from them.
- C- It changes the rate of the reaction.
- D- It does not undergo a chemical change.

3- When the temperature of the chemical reaction increases, the rate of the reaction increases due to the increase in

- A- The surface area exposed to the reaction.
- B- The number of reacting molecules.
- C- The number of possible collisions between molecules.
- D- The number of reacting atoms.

4- Enzymes work in many biological processes.

- A- As oxidizing agents
- B- As disinfecting agents
- C- As reducing agents
- D- As catalysts

Question two:

Complete the following sentences with suitable words:

- 1- The speed of chemical reactions increases by the increase in temperature.
- 2- Sweet Potatoes produce enzyme which increases the speed of the decomposition of
- 3- The catalyst changes the speed of the chemical reaction without affecting the or reaction.
- 4- In most modern cars have to treat harmful gases resulting from fuel combustion before they are expelled.

Question three: -

What are the results of: -

- 1- Replacing dilute HCl with concentrated acid when it reacts with magnesium strip.
- 2- Reaction of oil with caustic soda.